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# Water formation through $O_2 + D$ pathway on cold silicate and amorphous water ice surfaces of interstellar interest

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The formation of the first monolayer of water molecules on bare dust grains is of primary importance to understand the growth of the icy mantles that cover dust in the interstellar medium. In this work, we explore experimentally the formation of water molecules from  $O_2 + D$  reaction on bare silicate surfaces that simulates the grains present in the diffuse interstellar clouds at visual extinctions ( $A_V < 3$  mag). For comparison, we also study the formation of water molecules on surfaces covered with amorphous water ice representing the dense clouds ( $A_V \geq 3$  mag). Our studies focus on the formation of water molecules in the sub-monolayer and monolayer regimes using reflection absorption infrared spectroscopy and temperature-programmed desorption techniques. We provide the fractions of the products, such as  $D_2O$  and  $D_2O_2$  molecules formed on three astrophysically relevant surfaces held at 10 K (amorphous olivine-type silicate, porous amorphous water ice, and nonporous amorphous water ice). Our results showed that the formation of  $D_2O$  molecules occurs with an efficiency of about 55%–60% on nonporous amorphous water ice and about 18% on bare silicate grains surfaces. We explain the low efficiency of  $D_2O$  water formation on the silicate surfaces by the desorption upon formation of certain products once the reaction occurs between  $O_2$  and D atoms on the surface. A kinetic model taking into account the chemical desorption of newly formed water supports our conclusions. © 2012 American Institute of Physics. [<http://dx.doi.org/10.1063/1.4771663>]

## I. INTRODUCTION

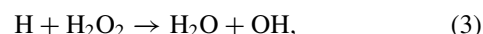
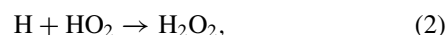
Solid water ice is the most abundant condensed component in the cosmos and particularly on the icy mantle of dust grains in dense molecular clouds. It is also the most abundant species in the icy mantles covering dust grains<sup>1</sup> in the interstellar medium (ISM). The accretion of water from the gas phase cannot explain the observed abundances of solid water in dark molecular clouds.<sup>2</sup> It was suggested in the past that water ice is formed through surface reactions where dust grains play a role of catalyst.<sup>3</sup> Three different hydrogenation routes for water formation on the surface of dust grains have been proposed by Tielens and Hagen:<sup>3</sup>  $O + H$ ,  $O_2 + H$ , and  $O_3 + H$ .

In dense molecular clouds, dust grains are covered with icy mantles which play the role of catalyst, and hydrogen is mainly in its molecular form. Hydrogen atoms are produced by cosmic rays with an abundance of  $(10^{-3} H_2)$ .<sup>4</sup>  $O_3$  and  $O_2$  molecules have not so far been detected in interstellar ices on dust grains,<sup>5</sup> while the abundance of O atoms in these environments, is predicted by astrochemical models to be  $(O/H = 10^{-4})$ .<sup>6</sup>

The  $O_3 + H$  channel is demonstrated theoretically<sup>7</sup> and experimentally<sup>5,8</sup> to be the most effective route for water formation in dense clouds.

According to Cuppen and Herbst,<sup>7</sup> the  $O_2 + H$  channel is also responsible for water formation on cold dense clouds. Several laboratory studies demonstrated the high efficiency ( $\sim 50\%$ ) of water ice formation through  $O_2$  and D atoms in the multilayer regime of oxygen ice at 10 K using reflection absorption infrared spectroscopy and temperature-programmed desorption (RAIRS)<sup>9–14</sup> and temperature-programmed desorption (TPD)<sup>11</sup> techniques.

The hydrogenation of molecular oxygen  $O_2$  results in the formation of water through the following exothermic reactions:



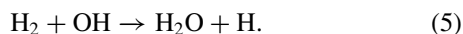
and



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Reaction (1) has a very low activation barrier of about 52 K in the gas phase,<sup>9</sup> while reaction (2) has no activation barrier.<sup>7,15</sup> However, despite the large activation barrier of the reaction (3) in the gas phase ( $\sim 1800$  K),<sup>14</sup> the experimental studies of Miyauchi *et al.*<sup>9</sup> at surface temperature of 10 K and Ioppolo *et al.*<sup>10,14</sup> at 12–28 K have confirmed the high efficiency of water formation through the  $O_2 + H/D$  pathway in the multilayer regime of  $O_2$  ice. Both  $H_2O_2$  ( $D_2O_2$ ) and  $H_2O$  ( $D_2O$ ) species were observed by RAIRS. Miyauchi *et al.*<sup>9</sup> and Oba *et al.*<sup>12</sup> have suggested a tunneling process for the reaction (3) because of the isotopic effect between  $H + H_2O_2$  and  $D + D_2O_2$  reaction rates. Recent experimental results of Cuppen *et al.*<sup>13</sup> showed that  $H_2O$  is not predominantly formed through the hydrogenation of  $H_2O_2$  but also through the reaction (4) between H atoms and OH radicals.

In dense clouds in which typical temperatures are about 10 K, the UV flux is very low, and  $H_2$  molecules the dominant H-bearing species.<sup>3,7,14</sup> Consequently, the reaction  $H_2 + OH$  (5) is often thought to be the most common route for the formation of  $H_2O$  on grain surfaces covered with water ices



Despite the high activation barrier of this reaction in the gas phase (2100 K),<sup>16</sup> Oba *et al.*<sup>17</sup> have demonstrated experimentally the formation of water ice at low surface temperature 10 K.

The formation of water by successive hydrogenation of atomic oxygen proceeds through the barrierless radical-radical reactions



Dulieu *et al.*<sup>18</sup> have demonstrated the high efficiency (50%) of water formation from O and D atoms on compact water ices substrates at 10 K using TPD experiments. Recently, Jing *et al.*<sup>19</sup> have studied experimentally the first stages of water formation through  $O + D$  reaction on the surface of an amorphous film of silicates at 15 K and 25 K using RAIRS and TPD techniques. They observed the formation of  $D_2O_2$ ,  $D_2O$ , and  $O_3$ , and suggested that the formation of  $D_2O_2$  molecules proceeds through  $OD + OD \rightarrow D_2O_2$  reaction because of the presence of the strong infrared absorption peak at  $2300\text{ cm}^{-1}$  of OD radicals on the surface.

Theoretical simulations (Cuppen and Herbst<sup>7</sup>) suggest that the  $O + H$  reaction pathway is the most dominant process for water formation in diffuse clouds in the ISM where bare dust grains are mostly composed of silicates and/or carbonaceous materials.<sup>20</sup> In such places gas phase composition is mostly atomic because of the high UV radiation fields. In these gas environments, where the visual extinctions ( $A_V < 1$  mag), the densities of hydrogen and oxygen atoms are expected to be high ( $n_H = 100\text{ atoms cm}^{-3}$ ) and ( $n_O = 0.03\text{ atoms cm}^{-3}$ ), with an atomic abundance of  $O/H = 10^{-4}$  (see Ref. 6).

Silicates are among the most commonly found dust species in the (ISM) of galaxies. They are often assumed to be metallic and dielectric materials.<sup>21</sup> Their presence is estab-

lished through the detection of the mid-infrared resonances due to the SiO stretching and the OSiO bending mode at 9.7 and  $18\text{ }\mu\text{m}$ , respectively.<sup>22</sup>

Recent gas-surface chemical models of Goumans *et al.*<sup>23</sup> showed that the formation of  $H_2O$  water ice on cold silicate surface such as forsterite can occur via species (O and H) strongly bound to the surface that can dissipate efficiently their excess energy to the surface via phonons. However, the exothermicity of the  $H + OH \rightarrow H_2O$  reaction could provide sufficient surplus energy to desorb  $H_2O$  water molecules into the gas phase from bare silicate surfaces. Certainly, the evaporation of water molecules at low surface temperatures (10 K) is inefficient, but their desorption into the gas phase upon formation on cold dust grains analogs, can be induced by the heat of formation released from the exothermic surface reaction  $H + OH$ . This desorption process of newly formed water species is called the chemical desorption,<sup>24</sup> it is also referred as non-thermal desorption as opposed to TPD thermal desorption.<sup>25</sup> This process has a great impact on the chemical composition of interstellar gas.

Recently, Cazaux *et al.*<sup>26</sup> studied water formation on dust grains using Monte Carlo simulations and provided the fractions of physisorbed molecules (OH,  $HO_2$ ,  $H_2O_2$ ,  $H_2O$ , OD,  $DO_2$ ,  $D_2O_2$ ,  $D_2O$ ) released into the gas phase upon formation in different astrophysical environments. These authors suggested that a significant part of the species are desorbed upon formation from bare surfaces because of the exothermicity of certain reactions on the surface. Then in diffuse clouds, 35% of OH desorbed upon formation from  $O + H$  reaction and 15% of  $H_2O$  is ejected in the gas phase from  $OH + H$  reaction  $H + OH$ .

In order to understand the growth of water ice mantle that covers the dust grains in dense molecular clouds, it is primordial to understand the process and the mechanism of water formation on the surface of bare dust grains composed either of silicates or carbonaceous materials. As previously mentioned, the formation of  $H_2O$  water ice in dense clouds is likely to occur from the hydrogenation of  $O_3$ ,  $O_2$ , and O atoms. However, the reactivity of oxygen atoms on the surface may form again oxygen and ozone molecules, respectively, from  $O+O$  and  $O+O_2$  reactions.<sup>19,27</sup> In this work, we investigate experimentally the formation of the first monolayer of  $D_2O$  water ice through  $O_2$  and D atoms on amorphous silicate surface held at 10 K. In order to study the effect of the substrate on the efficiency of water formation in the sub-monolayer and monolayer regimes, the experiments were also investigated both on porous and nonporous amorphous solid water ices.

The plan of this paper is the following: In Sec. II, we describe the experimental setup and we explain the experimental procedures. In Sec. III, we present the experimental results for the synthesis of water ice through  $O_2 + D$  channel on silicate and amorphous water ices substrates at 10 K. We used D atoms instead of H atoms to increase the signal to noise ratio in spectrometric measurements. Surface reactions are analyzed both with RAIRS and TPD spectroscopy. In Sec. IV, we discuss the results and we propose a reaction route for water formation, and a chemical model to support our conclusions.

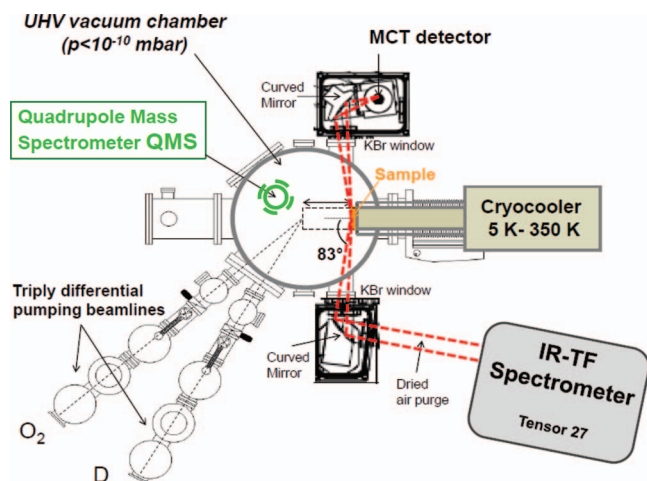


FIG. 1. Schematic of the experimental setup showing the UHV chamber, the atomic/molecular beam lines, the QMS, and the FTIR diagnostics.

## II. EXPERIMENTS

The experiments were performed using the formation of molecules in the ISM setup. It is briefly described here and more details are given in Refs. 28–31. The apparatus displayed in Figure 1 consists of an ultra-high vacuum (UHV) stainless steel chamber with a base pressure lower than  $1 \times 10^{-10}$  mbar. In its center, we find the sample holder, which is thermally connected to a cold finger of a closed-cycle He cryostat ARS cryogenics 210SD. The temperature of the sample is measured in the range of 5.5–350 K with a calibrated silicon diode clamped to the sample and connected to a Lakeshore 336 controller. The temperature is controlled within  $\pm 0.2$  K precision and  $\pm 1$  K of absolute accuracy depending on how tight is the silicon diode clamped to the surface.

The sample holder is made of a 1 cm diameter copper block. Its optically polished and gold-coated surface is covered with an amorphous olivine type silicate film  $[(\text{Mg}_x\text{Fe}_{1-x})_2\text{SiO}_4]$  where  $0 < x < 1$ . Its exact chemical composition is unknown, but its amorphous propriety is evidenced by infrared spectroscopic studies,<sup>32</sup> and its nonporous surface structure is estimated from TPD experiments.

An analytical quadrupole mass spectrometer (QMS Hi-den 3F) is used to quantify the reactants entering the chamber and the products desorbing from the sample surface. We used two triply differentially pumped beam-lines aimed at the sample to introduce the reactants ( $\text{O}_2$ , D) onto the surface of the sample holder. The two beam lines are equipped with micro-wave dissociation sources (a Surfatron cavity delivering 300 W at 2.45 GHz) that can generate deuterium atoms from their parent molecules. When the microwave sources are turned on, the dissociated efficiency of  $\text{D}_2$  beam is  $\tau = 65\%–70\%$ . The flux of D atoms hitting the surface is defined as  $\Phi_D = 2 \times \tau \times \Phi_{\text{D}_2}$  where  $\Phi_{\text{D}_2} = (2.2 \pm 0.4) \times 10^{12}$  molecules  $\text{cm}^{-2} \text{s}^{-1}$  is the flux of  $\text{D}_2$  molecules. The flux of D atoms is estimated to be  $(3.0 \pm 0.5) \times 10^{12}$  particles  $\text{cm}^{-2} \text{s}^{-1}$ , whereas the flux of  $\text{O}_2$  molecules at the surface is evaluated to  $(2.8 \pm 0.4) \times 10^{12}$  molecules  $\text{cm}^{-2} \text{s}^{-1}$ .

Since water molecules are polar and IR active, water ices are monitored *in situ* by means of RAIRS. RAIRS is a powerful surface infrared technique used to study the vibrational spectrum of very thin adsorbate on a surface.<sup>33</sup>

We used a Fourier transform infrared spectrometer (FTIR), Bruker Tensor 27 with a mid-IR source (Globar) emitting light between  $2.5 \mu\text{m}$  and  $25 \mu\text{m}$ . As shown in Figure 1, the spectrometer is coupled to the main UHV chamber through two differentially pumped potassium bromide (KBr) windows for the IR incident and reflected light. The spectrometer is used in external mode and the IR beam exiting it through a KBr window. Spectra are recorded in the  $4000–400 \text{ cm}^{-1}$  spectral region which includes absorption by several molecules detected in the ISM. As shown in Figure 1, the incoming IR beam is bent at  $90^\circ$  by a parabolic mirror with a focal distance  $f = 25 \text{ cm}$  before entering in the main chamber. This mirror focuses the IR beam on the sample at an incidence angle of  $\sim 83^\circ$  corresponding to a grazing angle of  $\sim 7^\circ$ . The beam reflected from the highly reflective surface is then focused at  $90^\circ$  by an ellipsoidal mirror ( $f = 25 \text{ cm}$ ) onto a mercury-cadmium-telluride (model D316/6) detector cooled with liquid nitrogen  $\text{N}_2$  (77 K). This detector is characterized by its fast response time and its high sensitivity. A spectral resolution of  $4 \text{ cm}^{-1}$  is used and 500 scans (about 400 s acquisition time) are co-added for one averaged spectrum in order to reduce the signal to noise ratio.

Our RAIRS experiments required two positions for the sample holder. During the exposure phase, the sample is placed in the center of the vacuum chamber where  $\text{O}_2$  and D beams are aimed at the surface. Then, the sample is moved back  $\sim 9 \text{ cm}$  from the exposure position for infrared measurements. Initially a spectrum of the clean surface at 10 K has been recorded before dosing which is called reference spectrum  $S_R$ . Then a sample spectrum  $S_S$  is acquired after exposure to  $\text{O}_2$  and D-atoms. The RAIR spectrum of the sample compound (resulting from the difference between  $S_S$  and  $S_R$ ) is plotted in absorbance (AB) mode for quantitative analysis.

Since the diatomic homonuclear molecule  $\text{O}_2$  is infrared inactive, gas phase  $\text{O}_2$  is monitored during the desorption by the QMS using TPD spectroscopy. It consists to heat the sample with a linear rate of heating  $dT/dt = 0.04 \text{ K s}^{-1}$  until the sublimation of water species.

The experiments of water formation from  $\text{O}_2$  and D-atoms has been performed at 10 K on different surface samples: (i) silicate (as described above), (ii) nonporous amorphous solid water (np-ASW) ice, and (iii) porous amorphous solid water (p-ASW) ice. The np-ASW ice film with a thickness of about 50 layers ( $1 \text{ ML} = 10^{15}$  molecules  $\text{cm}^{-2} \text{s}^{-1}$ ) is grown by spraying water vapor from a microchannel array doser located 2 cm in front of the surface held at 110 K. The np-water film (called also compact water ice) is then cooled down to 10 K for the exposure of  $\text{O}_2$  molecules and D atoms. Otherwise the p-ASW ice film with 10 layers of thickness is introduced by background vapor deposition on the compact water ice substrate kept at 10 K using the same microchannel array doser placed in the remote position.

The D,  $\text{O}_2$ , and  $\text{D}_2\text{O}$  beams have been calibrated to estimate the exposure time required for each gas to form the first monolayer coverage of the species on the surface of the



silicate at 10 K. The O<sub>2</sub> beam calibration consists of depositing different amounts of oxygen molecules on the surface and monitoring their thermal desorption by the QMS. The calibration of D-atoms beam was estimated from King and Wells<sup>34</sup> type experiments for the sticking coefficient of particle incident on the surface using mass spectrometry (see Ref. 29 for details about this measurement). In our experimental conditions, the 1 ML of D-atoms is evaluated to 7 min of D deposition time, the 1 ML of O<sub>2</sub> ice was obtained after 6 min of O<sub>2</sub> exposure time.

For all the experiments, the coverages (in ML) of the nascent O<sub>2</sub> and D species are estimated from the ratio between the exposure time of gas phase and exposure time required to form 1 ML coverage. For D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> species, the coverages (in ML) are obtained from the comparison of RAIRS integrated areas of the observed peaks and the integrated area corresponding to one monolayer coverage of D<sub>2</sub>O molecules. Because the coverages of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> were not possible by TPD, we have used the yield calculation. The yields (in %) of species (O<sub>2</sub>, D<sub>2</sub>O...) in solid phase are obtained by dividing the (RAIRS or TPD) integrated area of the corresponding peak, by the expected integrated area when O<sub>2</sub> is fully consumed by D atoms following the reaction  $\text{O}_2 + 4\text{D} \rightarrow 2\text{D}_2\text{O}$ .

### III. RESULTS

#### A. Formation of water on amorphous silicates

##### 1. RAIRS results

In these experiments, we exposed the surface of the silicate held at 10 K to many cycles of 1 min of O<sub>2</sub> + 4 min of D (i.e., 0.18 ML of O<sub>2</sub> + 0.58 ML of D), by increasing progressively the deposition times of O<sub>2</sub> and D atoms under the condition  $[\text{D}]/[\text{O}_2] \simeq 4$ ; it corresponds to the stoichiometric proportions of the reaction  $\text{O}_2 + 4\text{D} \rightarrow 2\text{D}_2\text{O}$ . This co-deposition method of O<sub>2</sub> and D atoms ensures that the reaction between O<sub>2</sub> and D atoms occurs in low-coverage regime

until all O<sub>2</sub> molecules consumed by D atoms are transformed into D<sub>2</sub>O molecules. We used this method of deposition to be able to observe the infrared signature of D<sub>2</sub>O water ice formed on the silicate surface at 10 K in the sub-monolayer coverages.

Despite the low sticking coefficient of D atoms ( $\sim 20\%$ ) on grain surfaces at 10 K,<sup>35</sup> the low fraction of the impinging D atoms that stuck on the surface react first with O<sub>2</sub> molecules and form DO<sub>2</sub> which can be deuterated into OD and D<sub>2</sub>O<sub>2</sub> species. The final products of the deuteration are D<sub>2</sub>O water molecules. The impinging D atoms can also react with another D atom on the surface and form D<sub>2</sub> molecules. The formation of hydrogen molecules occurs either via Eley-Rideal (E-R) mechanism between an adsorbed D atom and another one coming from the gas phase, or by Langmuir Hinshelwood mechanism which is enhanced by the diffusion of two adsorbed D atoms on the surface.<sup>7</sup>

In Figure 2 left panel, the RAIR spectra from (a) to (g) show the growth of the infrared absorption peaks of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> ices in the spectral region  $[3000\text{ cm}^{-1} - 1900\text{ cm}^{-1}]$  of the OD-stretching vibration bonds and in the spectral region  $[1300\text{ cm}^{-1} - 1000\text{ cm}^{-1}]$  of the OD-bending modes of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub>.

Because of the low sensitivity of our RAIRS diagnostic for low coverages, it was not obvious to observe clearly the IR absorption peaks of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> in the sub-monolayer regime (see spectra (a)–(c)). However, by increasing the exposure doses of O<sub>2</sub> and D atoms (1.1 ML O<sub>2</sub> + 4.2 ML D) to observe the monolayer regime (see spectrum (d)), we start observing a broad absorption peak at around  $2404\text{ cm}^{-1}$  corresponding to the overlapping of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> –OD stretching vibrations. The small peak at around  $2107\text{ cm}^{-1}$  is attributed to the  $2\nu_6$  vibration mode of D<sub>2</sub>O<sub>2</sub>.<sup>13,19</sup> In addition, a small peak at around  $1215\text{ cm}^{-1}$  is attributed to OD-bending mode of D<sub>2</sub>O and a second small peak at around  $1050\text{ cm}^{-1}$  is attributed to the OD-bending of D<sub>2</sub>O<sub>2</sub>.<sup>9,19</sup> RAIR spectra (e)–(g) in Figure 2 left panel show the growth of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub>

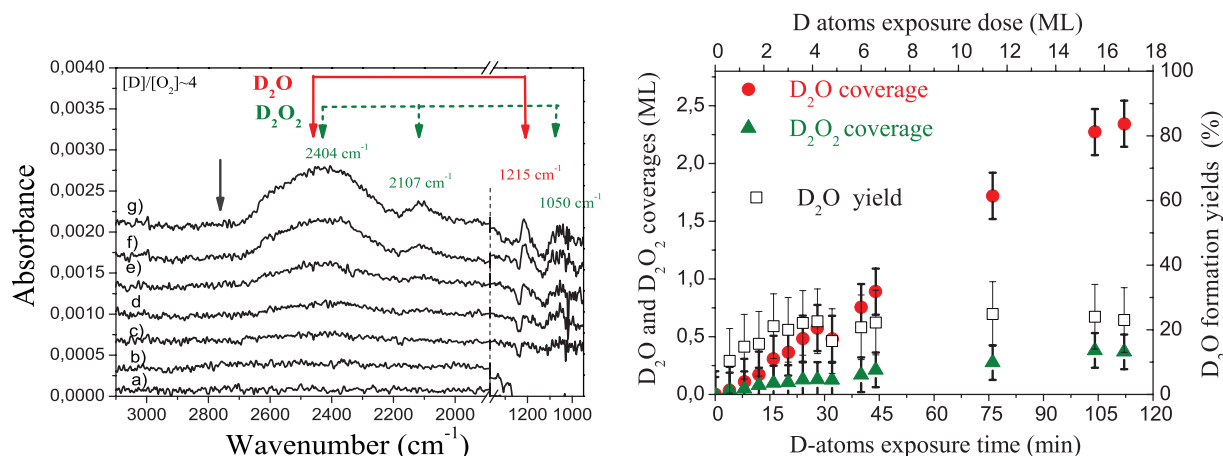


FIG. 2. (Left panel) RAIRS spectra (from (a) to (g)) of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> molecules formed on the silicate substrate held at 10 K by co-deposition of O<sub>2</sub> molecules and D atoms for several exposure times under the condition of O<sub>2</sub>/D = 4. (a) 0.2 ML(O<sub>2</sub>) + 0.6 ML(D), (b) 0.6 ML(O<sub>2</sub>) + 1.8 ML(D), (c) 0.7 ML(O<sub>2</sub>) + 2.4 ML(D), (d) 1.1 ML(O<sub>2</sub>) + 4.2 ML(D), (e) 2.0 ML(O<sub>2</sub>) + 6.6 ML(D), (f) 3.5 ML(O<sub>2</sub>) + 11 ML(D), and (g) 5.1 ML(O<sub>2</sub>) + 17 ML(D). (Right panel) Full circles and triangles, the estimated coverages (in ML) of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> on the silicate surface, respectively, as a function of D exposure doses and times. Empty blue squares, the yield of D<sub>2</sub>O molecules formed on the silicate surface, obtained from the formulae:  $f_{\text{D}_2\text{O}} = \frac{\text{D}_2\text{O coverage}}{2 \times \text{O}_2 \text{ coverage}}$ . The uncertainties are given by the errors bars.

peaks in the multilayer regime. We note that neither OD radicals having an absorption peak at about  $2300\text{ cm}^{-1}$  in solid phase,<sup>19</sup> nor  $\text{DO}_2$  radicals peaking at around  $(2549\text{ cm}^{-1})$ <sup>36</sup> were observed in our spectra during the co-deposition phase of  $\text{O}_2$  and D atoms on the silicate surface.

Figure 2 right panel shows the increase in the coverages (in ML) of  $\text{D}_2\text{O}$  and  $\text{D}_2\text{O}_2$  species on the silicate surface held at 10 K as a function of D-atoms exposure times and doses. In this experiment, the coverages are estimated from the infrared integrated areas of the small  $\text{D}_2\text{O}_2$  band ( $2\nu_6$ ) at  $2107\text{ cm}^{-1}$ , while the coverages of  $\text{D}_2\text{O}$  are estimated from the integrated areas of the wide band around  $2404\text{ cm}^{-1}$  of  $\text{D}_2\text{O}$  molecules overlapping  $\text{D}_2\text{O}_2$  species in the ( $\nu$  –OD) stretching mode. The combination of  $\text{D}_2\text{O}$  and  $\text{D}_2\text{O}_2$  modes in this spectral region leads to an over-estimation for  $\text{D}_2\text{O}$  formation. The ratio of the integrated areas for  $\text{D}_2\text{O}$  and  $\text{D}_2\text{O}_2$  peaks is found to be  $\text{D}_2\text{O}/\text{D}_2\text{O}_2 = \sim 4.0$  for very low surface coverages ( $\text{O}_2 \leq 1\text{ ML}$ ) and  $\text{D}_2\text{O}/\text{D}_2\text{O}_2 = \sim 6.5$  for ( $\text{O}_2 > 1\text{ ML}$ ).

The Figure 2 right panel shows that the amount of  $\text{D}_2\text{O}$  molecules is always higher than  $\text{D}_2\text{O}_2$  as the time of D atoms exposure increases. This is because in this co-deposition experiment, the monolayer fraction of  $\text{O}_2$  molecules deposited on the surface reacts each time with D atoms and forms mostly  $\text{D}_2\text{O}$  species. In the case of the sub-monolayer regime, the high abundance of  $\text{D}_2\text{O}$  on the surface in comparison to  $\text{D}_2\text{O}_2$  results from the diffusion of a sufficient number of D atoms on the surface of the silicate, favoring thus the conversion of most of OD and  $\text{D}_2\text{O}_2$  surface products into  $\text{D}_2\text{O}$  molecules. These results are different from those of Miyauchi *et al.*<sup>9</sup> and Ioppolo *et al.*<sup>10</sup> who observed much more  $\text{D}_2\text{O}_2$  than  $\text{D}_2\text{O}$  because of the diffusion of D(H) atoms in the multilayers of  $\text{O}_2$  ice. The  $\text{H}_2\text{O}$  ( $\text{D}_2\text{O}$ ) products formed on the top of hydrogen peroxide layers desorb into the gas phase by the heat released from the reactions.

The surface structure of the thin  $\text{D}_2\text{O}$  water ice formed on the silicate surface at 10 K from 5 ML of  $\text{O}_2$  and 17 ML of D atoms (RAIR spectrum (g), Figure 2 left panel) is not well identified. The absence of a small dangling (–OD) absorption feature around  $(2730\text{ cm}^{-1})$ <sup>37</sup> (see the arrow in the spectrum (g)) may be due to the high signal to noise ratio. However, the compact (not micro-porous) structure of the newly formed water ices from  $\text{O}_2 + \text{D}$  reaction has been largely studied and confirmed experimentally.<sup>12,38,39</sup>

Knowing the initial coverage of  $\text{O}_2$  molecules on the surface and that of  $\text{D}_2\text{O}$  molecules formed from D atoms and  $\text{O}_2$  molecules, we have estimated the yields of  $\text{D}_2\text{O}$  molecules formed on the surface as a function of D-atoms exposure times (see Figure 2 right panel). The yield of  $\text{D}_2\text{O}$  defined as the ratio between the amount of  $\text{D}_2\text{O}$  molecules formed on the surface and the amount of  $\text{D}_2\text{O}$  molecules expected to be formed from one molecule of  $\text{O}_2$  and four D atoms is given by the following formulae  $f_{\text{D}_2\text{O}} = \frac{\text{D}_2\text{O coverage}}{2 \times \text{O}_2 \text{ coverage}}$ . As shown in this figure, for the very low coverages of  $\text{O}_2$  and D atoms, the yield of  $\text{D}_2\text{O}$  is found to be lower than 15% probably because of the uncertainties of the infrared measurements. As soon as the coverage of water ice is about 0.5 ML, the formation yield of  $\text{D}_2\text{O}$  molecules stabilizes at about 20% and stagnates at this value for more than one monolayer coverage of  $\text{D}_2\text{O}$  ice.

It is possible that water molecules in the first monolayer do not cover uniformly the surface of the silicate because of the chemical desorption of the newly formed water molecules, which is enhanced by the metallic substrate. That is why the formation yield of water ice is low (20%). However, the deficiency of  $\text{D}_2\text{O}$  water molecules in some sites on the surface, may create small clustering water molecules<sup>26</sup> that may reduce the formation yield of water in the second and the subsequent layers.

## 2. TPD results

In these experiments, the surface of the silicate is uniformly covered with one monolayer coverage of solid  $\text{O}_2$  at 10 K to eliminate the effect the substrate, and then irradiated with D atoms for several exposure doses (0, 0.9 ML, 1.8 ML, 3.3 ML, and 5.5 ML). In these conditions, D atoms are likely to meet with an  $\text{O}_2$  molecule already adsorbed on the surface, react and form  $\text{O}_2\text{D}$ ,  $\text{D}_2\text{O}_2$ , OD, and  $\text{D}_2\text{O}$  species via Eley-Rideal mechanism. The competitive  $\text{D} + \text{D} \rightarrow \text{D}_2$  reaction with no activation barrier is always possible at low surface temperature,<sup>11</sup> but it is less probable at least as long as oxygen reactive species ( $\text{O}_2$ ,  $\text{DO}_2$ ,  $\text{D}_2\text{O}_2$ , and OD) are present on the surface. These oxygen populations consume D atoms via surface reactions with low energy barriers, reducing therefore the residence time of D atoms on the surface, and their density available for  $\text{D}_2$  formation.<sup>9–11,19</sup>

In our experiments, after each D-atoms exposure dose on 1 ML coverage of  $\text{O}_2$  ice film, the sample is then heated from 10 K to 220 K. Figures 3(a) and 3(b) left panel show the TPD curves of  $\text{O}_2$  between 10 K and 60 K,  $\text{D}_2\text{O}$  between 120 K and 180 K, and  $\text{D}_2\text{O}_2$  between 140 K and 220 K for different exposure doses of D atoms. By increasing the amount of D-atoms we see clearly the gradual destruction of  $\text{O}_2$  molecules adsorbed on the surface, and the growth of  $\text{D}_2\text{O}$  and  $\text{D}_2\text{O}_2$  ices. We point out that no intermediate species such as  $\text{O}_2\text{D}$  and OD radicals are detected by the QMS while heating the sample.

Figure 3 right panel shows the evolution of the yields (in %) of  $\text{O}_2$ ,  $\text{D}_2\text{O}$ , and  $\text{D}_2\text{O}_2$  species present on the surface as a function of D-atoms exposure doses on 1 ML coverage of  $\text{O}_2$  ice. These yields are estimated from the TPD integrated areas “a” of  $\text{O}_2$ ,  $\text{D}_2\text{O}$ , and  $\text{D}_2\text{O}_2$  peaks using the following formulae:  $f_{\text{O}_2} = 100 \times (\text{aO}_2/\text{aO}_2 \text{ consumed by D})$ ,  $f_{\text{D}_2\text{O}_2} = 100 \times (\text{aD}_2\text{O}_2/\text{aO}_2 \text{ consumed by D})$ , and  $f_{\text{D}_2\text{O}} = 100 \times (\text{aD}_2\text{O}/2 \times \text{aO}_2 \text{ consumed by D})$ . As shown in the Figure 3 right panel, after the exposure of  $\sim 1.5\text{ ML}$  of D atoms, about 85% of  $\text{O}_2$  molecules consumed by D atoms lead to the formation of  $\sim 19\%$  of  $\text{D}_2\text{O}_2$  and  $\sim 6\%$  of  $\text{D}_2\text{O}$  on the surface. As four D atoms are required in order to form two molecules of  $\text{D}_2\text{O}$ , a dose four times higher than 1.5 ML of D atoms (i.e.,  $\sim 4 \times 1.5\text{ ML} = 6\text{ ML}$ ) destroys all the  $\text{O}_2$  molecules present on the surface and forms only pure  $\text{D}_2\text{O}$  water ice with high efficiency (100%). However, our results show that by increasing the exposure of D atoms until 5.5 ML, the total amount of  $\text{O}_2$  molecules destroyed by D atoms reaches 98%, while the yield of  $\text{D}_2\text{O}$  molecules formed on the surface increases up to  $\sim 18\%$ .

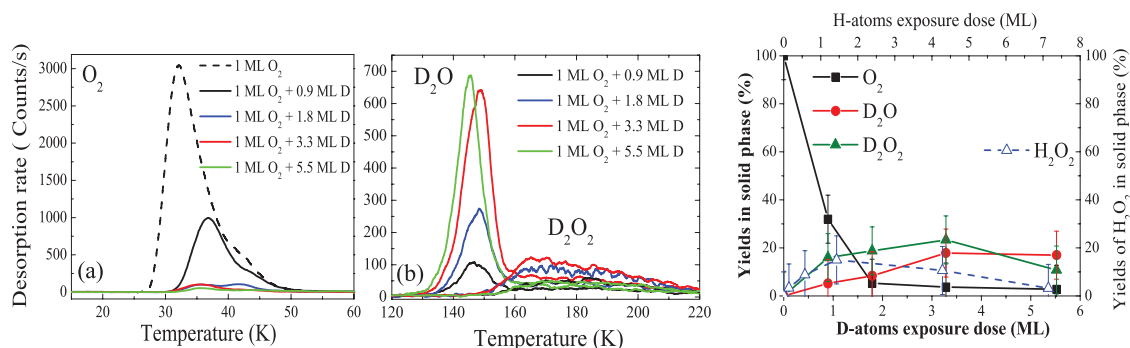


FIG. 3. (Left panel) (a) and (b) the desorption rates of  $O_2$ ,  $D_2O$ , and  $D_2O_2$  species as a function of D-atom exposure doses (0, 0.9 ML, 1.8 ML, 3.3 ML, and 5.5 ML) on 1 ML coverage of  $O_2$  ice film deposited on the silicate surface held at 10 K. (Right panel) The evolution of the yields (in %) of  $O_2$ ,  $D_2O$ , and  $D_2O_2$  species present on the surface of the silicate as a function of D-atom exposure doses. The blue dashed line is the yields (in %) of  $H_2O_2$  molecules present on the surface of the silicate as a function of H-atom exposure doses. The uncertainties of the measurements are marked by the errors bars.

For comparison, Figure 3 right panel displays the yields of  $H_2O_2$  obtained from  $O_2 + H$  reaction on the silicate surface held at 10 K for the same experimental conditions. The TPD data are plotted in dashed lines for clarity. We note that  $H_2O_2$  and  $D_2O_2$  have the same behavior for the low coverages of H(D) atoms on the surface ( $\leq 1.5$  ML). Then by increasing H(D)-atoms exposure doses, the  $H_2O_2 + H$  reaction seems to be faster than the  $D_2O_2 + D$  one thus suggesting a tunneling process for the reaction (3) on the silicate substrate as previously suggested by Miyauchi *et al.*<sup>9</sup> and Oba *et al.*<sup>12</sup>

Following these results, we suggest that the low amount ( $\sim 18\%$ ) of  $D_2O$  molecules formed on the surface of the silicate cannot be explained only by the tunneling process between D and  $D_2O_2$  species. We should take into account the possible desorption of certain products of the reaction such as  $DO_2$ ,  $D_2O_2$ ,  $OD$ , and  $D_2O$  once the reactions with D atoms occur on the surface. The results of Secs. III A 1 and III A 2 are summarized in Table I for the silicate.

TABLE I. Yields (in %) obtained from TPD and RAIRS experiments of  $O_2$  molecules consumed by D atoms,  $D_2O$ , and  $D_2O_2$  molecules formed from the  $O_2 + D$  reaction on different substrates held at 10 K: silicate SiOx, np-ASW ice, and p-ASW ice.

Experiments	TPD			RAIRS
	$T_{\text{surface}} \in [10 - 220 \text{ K}]$			$T_{\text{surface}} = 10 \text{ K}$
	$O_2$ ,	$D_2O_2$ ,	$D_2O$ ,	$D_2O$
Silicate				
1 ML $O_2$	0	0	0	...
1 ML $O_2 + 2.0 \text{ ML D}$	85	5	13	19
1 ML $O_2 + 3.3 \text{ ML D}$	96	24	18	...
1 ML $O_2 + 5.5 \text{ ML D}$	98	11	17	...
np-ASW ice				
1 ML $O_2 + 2.0 \text{ ML D}$	79	1	32	63
1 ML $O_2 + 3.3 \text{ ML D}$	85	5	55	...
p-ASW ice				
1 ML $O_2 + 2.0 \text{ ML D}$	60	0.1	28	45

## B. Effect of the substrate on the efficiency of the $O_2 + D$ reaction

In Sec. VI, RAIRS results showed that the yield of  $D_2O$  water molecules formed from  $O_2$  and D atoms on the silicate substrate held at 10 K reaches a steady state of  $\sim 20\%$  regime once the  $D_2O$  water ice coverage grows up to  $\sim 0.5$  ML. Such a result is supported by TPD experiments and stresses the ability of the silicate surfaces to desorb newly formed water species into the gas phase even for low surface coverages. In order to confirm these observations related to bare silicate grains, we have investigated the formation of the first  $D_2O$  water molecules from 1 ML coverage of  $O_2$  and 2 ML of D atoms on two substrates with different compositions (bare silicate and water ices). In dark interstellar clouds, dust grains are covered with icy mantles mainly composed of amorphous water ice whose morphology still remains poorly known. So we have also studied the formation of  $D_2O$  water molecules on two types of amorphous water ices surfaces (porous amorphous water ice with a large effective surface area available for the adsorption and the nonporous (or compact) amorphous water ice). Thanks to the mobility of D atoms on the surface at 10 K, the formation of  $D_2O$  water ice via  $O_2 + D$  reaction has been already observed experimentally in the submonolayer regime on porous water ice surfaces by Matar *et al.*<sup>11</sup>

### 1. RAIRS results

Figure 4 displays the RAIR spectra recorded in the  $-OD$  stretching region after the exposure of 2.0 ML of D atoms on 1 ML coverage of  $O_2$  ice film adsorbed on three substrates held at 10 K: amorphous silicate (spectrum (a)), p-ASW ice film with 10 ML of thickness (spectrum (b)), and np-ASW ice film with 50 ML of thickness (spectrum (c)). Here, the ratio D/ $O_2$  between the deposition of D atoms and  $O_2$  molecules is taken to be 2 and the flux of D atoms is  $3.0 \times 10^{12} \text{ atoms cm}^{-2} \text{ s}^{-1}$ .

As shown in Figure 4, the three spectra show a broad absorption band at around  $2400 \text{ cm}^{-1}$  attributed to  $D_2O$  water ice overlapping  $D_2O_2$  molecules. The strong absorption bands observed in spectra (b) and (c) are an indication that



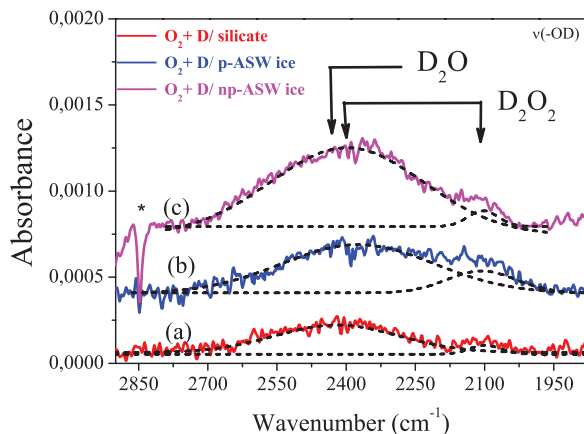


FIG. 4. RARS spectra of  $D_2O$  and  $D_2O_2$  ices formed after the exposure of 1 ML of solid  $O_2$  and 2.0 ML of D atoms on different substrates held at 10 K: (a) amorphous silicate, (b) p-ASW ice, and (c) np-ASW ice. The dashed curves are the Gaussian fits. (The star at  $2850\text{ cm}^{-1}$  is an impurity.)

the amounts of  $D_2O$  molecules formed both on p-ASW and np-ASW ices are higher than that formed on bare silicate substrate. The shoulder at  $2105\text{ cm}^{-1}$  is clear evidence for  $D_2O_2$  molecules formed from the  $O_2 + D$  reaction.

Since  $D_2O$  band overlaps with  $D_2O_2$ , a multi-Gaussian fit is used to separate the contributions of  $D_2O$  at  $2400\text{ cm}^{-1}$  and  $D_2O_2$  at  $2105\text{ cm}^{-1}$ , and to determine the area of the individual bands. The proportions between the different integrated area of  $D_2O$  bands are found to be 3.5/2.3/1, for np-ASW, p-ASW, and silicate, respectively. By dividing the area of the broad absorption bands at  $2400\text{ cm}^{-1}$  by the integrated area of  $D_2O$  water peak corresponding to 1 ML coverage (not shown in Figure 4), we estimated the amount of  $D_2O$  water molecules formed on each substrate held at 10 K to be  $\sim 0.37$  ML on the silicate surface,  $\sim 0.89$  ML on the p-ASW ice, and  $\sim 1.25$  ML on np-ASW ice surfaces. These results give preliminary estimations for the yields of the species formed on the surface. It appears that  $\sim 19\%$  of the newly formed  $D_2O$  molecules remain on the silicate surface, while  $\sim 45\%$  of  $D_2O$  molecules reside on the p-ASW ice, and  $\sim 63\%$  on the np-ASW ice surfaces (see Table I).

## 2. TPD results

After the exposure of 2.0 ML of D atoms at 10 K on one monolayer coverage of  $O_2$  ice film, the sample is then heated from 10 K to 220 K until the complete desorption of the species implicated in the  $O_2 + D$  reaction. Figures 5(a) and 5(b) display for three substrates (silicate, p-ASW ice, and np-ASW ice films) the TPD curves of  $O_2$ ,  $D_2O$ , and  $D_2O_2$  molecules.

By comparing TPD curves of solid  $O_2$  before and after D atoms irradiation for different substrates, we can assume from Figure 5(a) that  $\sim 85\%$  of  $O_2$  molecules adsorbed both on silicate and np-ASW ice film are consumed by D atoms, while  $\sim 60\%$  of  $O_2$  molecules present on the p-ASW ice are destroyed after the exposure of 2.0 ML of D atoms. The high amount ( $\sim 40\%$ ) of  $O_2$  molecules not struck by D atoms on the porous water ice can be explained by the lack of some D

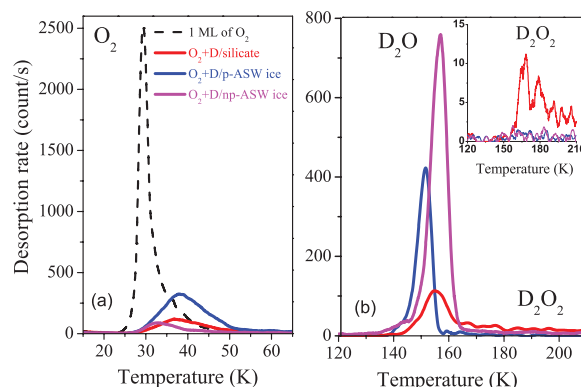


FIG. 5. Desorption peaks of  $D_2O$ ,  $D_2O_2$ , and  $O_2$  obtained after the exposure of 1 ML of  $O_2$  and 2.0 ML of D atoms coverages on different substrates held at 10 K: red color (silicate), blue color (p-ASW ice), and magenta (np-ASW ice). The inset shows the TPD curves of  $D_2O_2$  formed on three substrates.

atoms on the surface because of their diffusion at 10 K in the micro-pores structures of the water ice film.<sup>11</sup>

In addition, Figure 5(b) displays the desorption peaks of  $D_2O$  molecules formed on three substrates. The blue and the magenta peaks of  $D_2O$  molecules are corrected from the contribution of HDO molecules (not shown in the figure), produced by isotopic exchange during heating sample between the newly formed  $D_2O$  molecules and the  $H_2O$  molecules of water substrates.<sup>18</sup> The inset in Figure 5(b) shows a small desorption peak of  $D_2O_2$  molecules between 160 K and 210 K mainly for the silicate substrate. While the RAIR spectra in Figure 4 revealed the presence of  $D_2O_2$  peak at  $2107\text{ cm}^{-1}$  for the three substrates at 10 K (spectra (a)–(c)), the  $D_2O_2$  products formed on each substrate are not converted to  $D_2O$  water ice either by lack of D atoms, or by the presence of a high activation barrier for the  $D_2O_2 + D$  reaction. However, the quasi absence of  $D_2O_2$  desorption peaks both for p-ASW and np-ASW ices (see Figure 5(b)) can be explained by the possible isotopic exchange between the newly formed hydrogen peroxide  $D_2O_2$  and the  $H_2O$  molecules of water substrates during the heating phase of the sample. In fact, some other isotopic species than HDO ( $m/z = 19$ ) have been observed as traces in our TPD experiments mainly with the porous and the nonporous water ices substrates, such as  $HDO_2$  ( $m/z = 35$ ) and  $H_2O_2$  ( $m/z = 34$ ), but their desorption peaks are not shown in Figure 5(b).

Knowing the TPD integrated area of the newly formed  $D_2O$  molecules and that of  $O_2$  molecules consumed by D atoms, we have estimated the yields of  $D_2O$  molecules formed on each substrate to be 32% on np-ASW ice, 28% on p-ASW ice, and only 13% on the silicate substrate. These estimations are slightly lower than those previously obtained by RARS measurements.

Further TPD experiments were performed on the silicate and np-ASW ice substrates held at 10 K by lowering the flux of D-atoms to  $2.5 \times 10^{12}\text{ molecules cm}^{-2}\text{ s}^{-1}$  and increasing the D-atoms exposure doses up to 3.3 ML on one monolayer coverage of  $O_2$ . In this case, we observe much more  $D_2O_2$  not converted into water ice on silicate substrate compared to the previous experiments performed with higher D-atom flux.

The TPD results showed the formation of 18% of  $D_2O$  and 24% of  $D_2O_2$  on the silicate substrate at 10 K when most of  $O_2$  molecules (96%) are consumed by D-atoms. While in the case of the np-ASW ice substrate, the yield of  $D_2O$  molecules reaches its maximum of about 55% and that of  $D_2O_2$  5% when 85% of  $O_2$  molecules are destroyed by D-atoms. Table I summarizes for the different substrates (silicate, np-ASW ice, and p-ASW ice) held at 10 K the yields of  $D_2O$ ,  $D_2O_2$ , and  $O_2$  species present on the surface after D-atoms exposure on 1 ML coverage of  $O_2$  ice film. Table I shows that the TPD yields for  $D_2O$  water formation for the same experiments are always lower than the RAIRS ones because of the upper limits of water formation abundances by RAIRS measurements as previously mentioned in Sec. III A 1.

As previously stated, the combination of  $D_2O$  and  $D_2O_2$  absorption modes in the  $-OD$  stretching region leads to an over-estimation of the RAIR water formation yields. Table I also shows that for the same TPD experiments, the very low yields of  $D_2O_2$  for p- and np-ASW ices compared to silicate substrate, may result from an isotopic exchange between  $D_2O_2$  and  $H_2O$  molecules of water ices substrates.

In addition, the low formation yields of  $D_2O$  by TPD experiments in comparison to the yields measured by RAIRS (see Table I), may be explained by the isotopic exchange between  $D_2O$  and  $H_2O$  water molecules of porous and non-porous ices between (140 and 160 K), leading to the formation of HDO species ( $D_2O + H_2O \rightarrow 2HDO$ ).

### C. The chemical desorption process of water

We have experimentally observed the desorption upon formation of  $D_2O$  molecules from the silicate substrate held at 10 K once the reaction occurs between  $O_2$  and D atoms. Figure 6 shows the signal of  $D_2O$  ( $m/z = 20$ ) as well as the signals of OD ( $m/z = 18$ ) and  $D_2$  ( $m/z = 4$ ) species monitored in the gas phase with the QMS (placed very close to the sample) during 37 min (2220 s) of D-atoms on 1 ML of  $O_2$  ice.

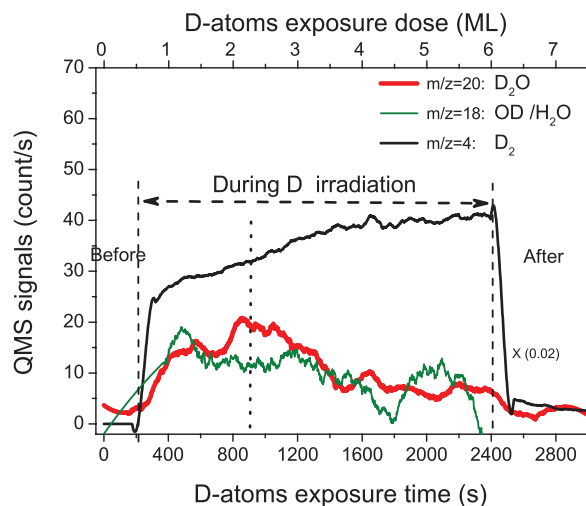


FIG. 6. The signals of  $D_2O$ , OD, and  $D_2$  species monitored by the QMS in the gas phase during 37 min of D-atoms exposure on 1 ML coverage of  $O_2$  ice deposited on the silicate surface held at 10 K.

The signal of mass ( $m/z = 18$ ) is noisy and can be attributed either to OD products released into the gas phase and/or  $H_2O$  molecules present as an impurity in the vacuum chamber. Figure 6 also shows that during the first 1000 s of D-atoms exposure, the amount of the newly formed  $D_2O$  molecules which are released from the surface into the gas phase increases and reaches a maximum of desorption, meaning that the formation of  $D_2O$  water molecules occurs efficiently on the surface of the silicate in the beginning of the D irradiation phase, and probably through the barrierless  $D + OD \rightarrow D_2O$  reaction. Afterwards, by increasing the D-atoms irradiation time until 2200 s, the signal of the newly formed  $D_2O$  molecules desorbing into the gas phase decreases gradually as the surface is depleted in  $O_2$  molecules. We note that no desorption into the gas phase of  $DO_2$  and  $D_2O_2$  species was observed in this experiment during all the D-atoms exposure phase.

In parallel, Figure 6 also shows the increase in the signal of  $D_2$  species during the exposure phase of D atoms, which corresponds to the desorption into the gas phase of  $D_2$  molecules. The linear increase of  $D_2$  molecules in the gas phase is enhanced by the increase of the coverage of D atoms and  $D_2$  molecules coming either from the fraction (30%) of the non-dissociated deuterium molecules from the D atomic beam, or from the fraction of D atoms that recombine with other D atoms on the surface of the silicates and form  $D_2$  by the exothermic  $D + D$  reaction. It is difficult to estimate the contribution of  $D_2$  molecules due exclusively to  $D + D$  reaction taking place on the sample, since D atoms are also consumed by the oxygen species ( $O_2$ ,  $DO_2$ , OD,  $D_2O_2$ ) present on the surface forming thus  $D_2O$  ice which is released from the surface by chemical desorption. Even if  $D_2$  are formed on the surface at 10 K, the local energy released from the  $D + D$  reaction does not affect the desorption of the newly formed  $D_2O$  and the other products.

The  $D_2$  species that are stuck on the surface at 10 K do not react with  $O_2$  molecules already adsorbed on the surface because of the very high activation barrier of the reaction.<sup>26</sup> The absence of any product when  $D_2$  beam is exposed to  $O_2$  ice has been already confirmed experimentally by Miyauchi *et al.*<sup>9</sup> and Cuppen *et al.*<sup>13</sup>

Moreover, Cuppen *et al.*<sup>13</sup> have suggested in their discussions using hydrogen instead of deuterium that some  $H_2$  molecules present on the surface may react with reactive species not desorbed into the gas phase such as  $HO_2$  and OH radicals and form  $H_2O_2$  and  $H_2O$ , respectively, through the following reactions  $H_2 + HO_2 \rightarrow H_2O_2 + H$  and  $H_2 + OH \rightarrow H_2O + H$ . But the efficiency of these reactions were below the detection limit.

### IV. ANALYSIS AND DISCUSSIONS

RAIRS and TPD results show that the formation of water molecules on cold surfaces through  $O_2 + D$  pathway is efficient in the sub-monolayer and monolayer regimes, but the fraction of the newly formed  $D_2O$  molecules which remains in the solid phase on the surface depends significantly on the substrate (see Table I).

In the case of the silicate substrate, the fraction of D<sub>2</sub>O water molecules that stays on the surface at 10 K upon formation through O<sub>2</sub> + D reaction is about 17% of the total amount of D<sub>2</sub>O molecules expected to be formed on the surface from 1 ML of O<sub>2</sub> and 5.5 ML of D-atoms (see Table I). As previously mentioned in Sec. III C, the lack of D<sub>2</sub>O molecules (~83%) on the surface of the silicate results from the desorption into the gas phase of newly formed D<sub>2</sub>O molecules.

Many factors enhance the desorption of species from a substrate: the surface roughness, the binding energies of the final products in the physisorbed sites of the surface, and mainly the energy of formation released from the exothermic reactions.

We have estimated the binding energies of D<sub>2</sub>O and D<sub>2</sub>O<sub>2</sub> molecules from our TPD experiments using the Polanyi-Wigner equation<sup>40</sup>

$$-\frac{dN}{dt} = r = AN^n \exp(-E_d/k_B T), \quad (8)$$

where  $r$  is the rate of desorption,  $N$  is the number of molecules adsorbed,  $n$  is the order of reaction,  $E_d$  is the desorption energy,  $k_B$  is the Boltzmann constant, and  $T$  is the absolute temperature of the surface. Assuming  $n = 0$  for water desorption in the sub-monolayer regime, the pre-exponential factor  $A$  is taken to be  $10^{13} \text{ s}^{-1}$ . The binding energies of species adsorbed on the silicate surface are estimated to be  $(465 \pm 10) \text{ meV}$  for D<sub>2</sub>O and  $(525 \pm 20) \text{ meV}$  for D<sub>2</sub>O<sub>2</sub> molecules and they are in agreement with the estimations of Jing *et al.*<sup>19</sup> on amorphous silicates held at 15 K. These values are also very close to the binding energies obtained from our estimations on nonporous amorphous water ice substrate for D<sub>2</sub>O ( $480 \pm 10$ ) meV and D<sub>2</sub>O<sub>2</sub> ( $520 \pm 40$ ) meV confirming thus the compact structure of our silicate surface.

In the case of the amorphous water ice substrates, the process of desorption upon formation is less significant than previously mentioned on bare grain silicates. It is possible that the catalytic effect of the water ice substrate enhances the reaction rates between D and O<sub>2</sub> by increasing the binding energies and reducing the exothermicity of the reactions. Such a catalytic effect of amorphous water ice has been demonstrated experimentally and theoretically by Hidaka, Kouchi, and Watanabe<sup>41</sup> for the successive H + CO<sub>2</sub> reactions. As a result of this effect, the newly formed species (O<sub>2</sub>D, OD, D<sub>2</sub>O<sub>2</sub>, and D<sub>2</sub>O) are likely to remain trapped on the surface sites of the amorphous water ice and dissipate their excess energy of formation on the surface until the thermalization.<sup>26</sup> These assumptions are in agreement with Jones and Williams<sup>42</sup> who suggested that the product of the reactions O + H → OH and OH + H → H<sub>2</sub>O should be retained on the grains with a probability of 70% and 100% for grain-mantle formation in dense clouds.

### A. Possible reactions routes for water formation

We demonstrated experimentally that the efficiency of water formation through O<sub>2</sub> + D pathway depends significantly on the substrate (silicates or amorphous water ice) held at 10 K. Figure 7 illustrates the possible reaction routes for

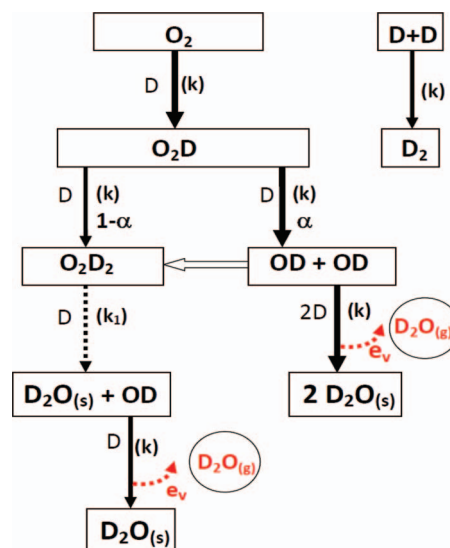
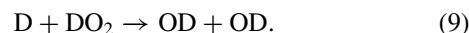


FIG. 7. Schematic illustration showing the possible reaction routes for the formation of solid D<sub>2</sub>O involving O<sub>2</sub> and D atoms. The solid and the broken black arrows indicate the reactions without and with activation barriers, respectively. Thick arrows denote reactions investigated in the present study. White arrow indicates the reaction which is less probable on bare silicate substrate than on amorphous water ice. D<sub>2</sub>O molecules desorbed in the gas phase upon formation are enclosed in circles.  $k$  is the rate constants of all the reactions without activation barrier and  $k_1$  the rate constant of the reaction with barrier.  $\alpha$  is the branching ratio of the reaction D + DO<sub>2</sub>, and  $e_v$  denotes the evaporation of D<sub>2</sub>O into the gas phase upon formation.

the formation of solid D<sub>2</sub>O water ice on the silicate substrate involving O<sub>2</sub> and D atoms. As shown in Figure 7, the impinging D atoms are likely to react via Eley-Rideal mechanism with oxygen O<sub>2</sub> molecules adsorbed on the surface or with the newly formed species DO<sub>2</sub>, D<sub>2</sub>O<sub>2</sub>, and OD. All these reactions are in competition with the D + D → D<sub>2</sub> reaction leading to the formation of D<sub>2</sub> molecules either via E-R mechanism between a D atom already adsorbed on the surface and an impinging D atom coming from the gas, or Langmuir-Hinshelwood mechanism between two adsorbed D-atoms on the surface. D atoms start to react with O<sub>2</sub> molecules with one monolayer coverage and form DO<sub>2</sub> products. The reaction involving DO<sub>2</sub> and D atoms may proceed through two alternative pathways. The first pathway D + DO<sub>2</sub> → 2OD (9) leading to 2OD intermediates has been recently highlighted by Cuppen *et al.*<sup>13</sup> for the hydrogen species

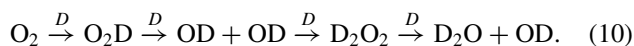


This exothermic reaction with an enthalpy  $\Delta H = \sim -1500 \text{ meV}$  seems to be predominant over the second classical pathway D + DO<sub>2</sub> → D<sub>2</sub>O<sub>2</sub> with higher activation barrier ( $\Delta H = -3540 \text{ meV}$ ), because it favors both the formation of water ice through OD + D and its chemical desorption into the gas phase. If the two OD radicals are formed in pairs in the same reaction,<sup>13</sup> they can react together because they are very close to each other, and form D<sub>2</sub>O<sub>2</sub> molecules on the surface following this reaction OD + OD → D<sub>2</sub>O<sub>2</sub> which has been recently evidenced by Cuppen *et al.*<sup>13</sup> and Jing *et al.*<sup>19</sup> Otherwise, the two OD intermediates react with D atoms and form D<sub>2</sub>O water molecules through the D + OD → D<sub>2</sub>O reaction without activation

barrier. We note that in the case of the silicate substrate, the two OD radicals are likely to be formed separately, and their association via Langmuir-Hinshelwood mechanism is difficult because of their low mobility at 10 K. This situation is different on amorphous water ice substrate, where the diffusion of the OD intermediates at low surface temperature can be enhanced by the hydrogen bonds of the nearest water molecules composing the substrate.

So most of  $D_2O$  molecules formed on the silicate surface from the barrierless  $D + OD$  reaction are likely to desorb into the gas phase upon formation because of the sufficient heat released from the exothermic reaction ( $\Delta H = -4290$  meV). The formation of  $D_2O$  water molecules through  $D + D_2O_2 \rightarrow D_2O + OD$  reaction is also possible but less probable because of the high activation barrier ( $\sim 2000$  K)<sup>17</sup> of the reaction. However, the amount of  $D_2O$  molecules formed slowly by quantum tunneling process during the D atoms irradiation is likely to remain on the surface, while some of OD radicals may react with D atoms coming from the gas phase or already adsorbed on the surface and form  $D_2O$  molecules without barrier. This scenario is in agreement with our previous TPD results concerning the slow evolution of  $D_2O_2$  and  $D_2O$  species on the surface with the increase of D atoms exposure on 1 ML of  $O_2$  ice (see Sec. III A 2, Figure 3 right panel).

In the case of the amorphous water ice substrates held at 10 K (porous and nonporous water ices), additional reaction routes can be suggested in order to explain the high percentages (42%–60%) of  $D_2O$  water formation on these surfaces through  $O_2 + D$  channel. In particular, the subsequent reaction  $OD + OD \rightarrow D_2O_2$  seems to be non-negligible on the amorphous water ice, and can contribute efficiently to the formation of  $D_2O$  water molecules on the surface through the following deuteration reactions:



Some OD radicals either react with D atoms to form  $D_2O$  molecules through the barrierless  $OD + D$  reaction (see Figure 7), or with some  $D_2$  molecules present on the surface and form new  $D_2O$  molecules via reaction (11) with high activation barrier of  $\sim 2100$  K<sup>16,17</sup>



## B. Modeling the $O_2 + D$ reaction on the silicates

In order to confirm our previous suggestions concerning the effect of the desorption of the newly formed oxygen species on the efficiency of water formation through  $O_2 + D$  pathway, a simple kinetic chemical model was developed to reproduce our TPD experimental results for the reaction of  $O_2$  with D on the silicate substrate held at 10 K. The model is made to fit the amounts  $[O_2]$ ,  $[D_2O_2]$ , and  $[D_2O]$  (in ML) of  $O_2$ ,  $D_2O_2$ , and  $D_2O$  species present on the surface of the silicate sample as a function of D-atoms exposure doses (0, 0.9 ML, 1.8 ML, 3.3 ML, and 5.5 ML) on 1 ML coverage of  $O_2$  ice film (see Sec. III A 2 and Figure 3 left panel).

Contrary to  $O_2$  and  $D_2O$  species, we were not able to measure by TPD experiments the coverage in (ML) of  $D_2O_2$

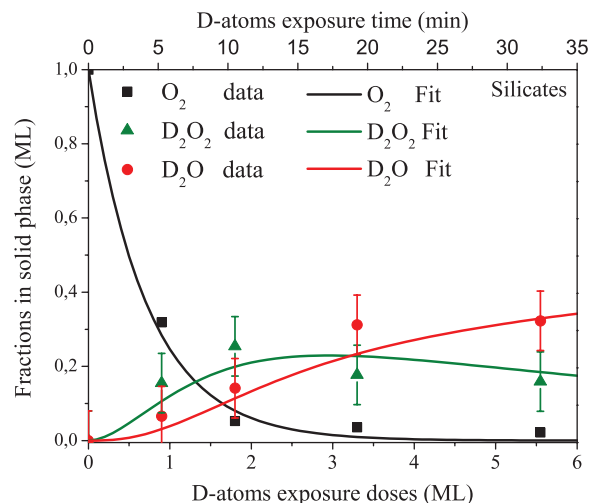
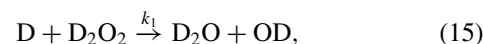
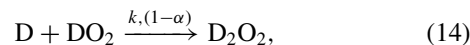
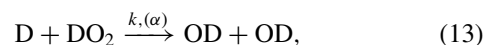


FIG. 8. (Full triangles, circles, and squares) The amount (in ML) of  $O_2$ ,  $D_2O$ , and  $D_2O_2$  species present on the surface as a function of D-atoms exposure doses on 1 ML coverage of  $O_2$  ice deposited on the silicate sample at 10 K. (Solid lines) The fits of  $O_2$ ,  $D_2O$ , and  $D_2O_2$  obtained from the model with  $\alpha = 0.69$ ,  $k = 1$ ,  $k_1 = 0.09$ ,  $\mu = 1$ ,  $\mu_1 = 0.14$ , and the fraction of  $D_2O$  desorbed from upon formation by the exothermic  $OD + D$  reaction is  $e_v = 86\%$ . The uncertainties are given by the errors bars which are equal to 0.08 for  $D_2O$  and 0.1 for  $D_2O_2$ .

ice on the surface, the calibration of  $D_2O_2$  beam was not possible. We then estimated the fractions in monolayer of the species present on the surface by normalization to the amount of  $O_2$  in 1 ML exposure coverage, see the following formulae:  $[O_2] = \text{area } O_2 / (\text{area of 1 ML } O_2)$ ,  $[D_2O_2] = \text{area } D_2O_2 / (\text{area of 1 ML } O_2)$ , and  $[D_2O] = \text{area } D_2O / (\text{area of 1 ML } O_2)$ . The experimental data are plotted in Figure 8.

The model is described by the following reaction routes which are illustrated in Figure 7 of Sec. IV A



The rate equations for the amount of  $O_2$ ,  $O_2D$ ,  $D_2O_2$ ,  $OD$ , and  $D_2O$  species present on the surface are given by

$$\frac{d[O_2]}{dt} = -n_D k [O_2], \quad (17)$$

$$\frac{d[O_2D]}{dt} = n_D (\mu k [O_2] - k [O_2D]), \quad (18)$$



$$\frac{d[\text{O}_2\text{D}_2]}{dt} = n_D [(1 - \alpha)\mu k[\text{O}_2\text{D}] - k_1[\text{O}_2\text{D}_2]], \quad (19)$$

$$\frac{d[\text{D}_2\text{O}]}{dt} = n_D(\mu k_1[\text{D}_2\text{O}_2] + \mu_1 k[\text{OD}]), \quad (20)$$

$$\frac{d[\text{OD}]}{dt} = n_D(2\alpha\mu k[\text{O}_2\text{D}] + \mu k_1[\text{O}_2\text{D}_2] - k[\text{OD}]). \quad (21)$$

We assume that D atoms coming from the gas phase react via Eley-Rideal mechanism with oxygen molecules covering the sample or with molecules previously formed. Where  $n_D = 2\tau\phi = 3.4 \times 10^{12} \text{ molecules cm}^{-2} \text{ s}^{-1}$  is the number of D-atoms per second coming from the gas phase and hitting the surface ( $\phi$  is the flux of deuterium molecules and  $\tau = 70\%$  is the dissociation rate of the  $\text{D}_2$  beam).  $k$  is the rate constant of the reactions (12)–(14) and (16), which are supposed to take place without a barrier.  $k_1$  is the rate constant of Eq. (15) with activation barrier. The  $\mu$  and  $\mu_1$  coefficients are the fractions of just formed species, for each reaction, which stay adsorbed on the sample.  $\mu$  corresponds to the fraction of the newly formed species through the reactions (12), (13), and (15), and  $\mu_1$  that of  $\text{D}_2\text{O}$  formed through the reaction (16). In order to keep the model as simple as possible we assumed  $\mu = 1$  and  $\mu_1 = 0.14$  from the experiments, so 86% of  $\text{D}_2\text{O}$  molecules desorbs into the gas phase upon formation. The idea was that the main contribution to the desorption is due to  $\text{D}_2\text{O}$  formed through reaction  $\text{OD} + \text{D}$  (16). Reaction (13) may proceed through two alternative pathways, and  $\alpha$  is the branching ratio (see Figure 7).

In our model, we assumed  $k = 1$  and fitted our data using  $\alpha$  and  $k_1$  as free parameters. Figure 8 shows the kinetic curves obtained from the model (plotted in solid lines) for  $\text{O}_2$ ,  $\text{D}_2\text{O}_2$ , and  $\text{D}_2\text{O}$ . The best fit is obtained with  $\alpha = 0.69$  and  $k_1 = 0.09$ , and it is always better for  $\text{D}_2\text{O}$  than  $\text{D}_2\text{O}_2$  because of the high uncertainty in the TPD measurements of hydrogen peroxide peak. The ratio of the reaction rate constants  $k_1/k = 0.09$  for  $\text{D} + \text{O}_2$  and  $\text{D} + \text{D}_2\text{O}_2$  reactions is found to have the same order of magnitude as that obtained (0.04) by the gas-surface model of Miyauchi *et al.*<sup>9</sup>

The model shows that the evolution of  $\text{O}_2$  molecules on the surface follows an exponential decay, and that  $\sim 95\%$  of  $\text{O}_2$  molecules destroyed by  $\sim 1.8 \text{ ML}$  of D-atoms produce about 0.1 ML of  $\text{D}_2\text{O}$  ice and 0.2 ML of  $\text{D}_2\text{O}_2$ . The presence of  $\text{D}_2\text{O}_2$  species on the surface even after a high exposure dose (5.5 ML) of D-atoms is an indication that the conversion of  $\text{D}_2\text{O}_2$  into  $\text{D}_2\text{O}$  molecules occurs slowly through reaction (15) which proceeds through tunneling process. Although the reaction  $\text{D} + \text{OD} \rightarrow \text{D}_2\text{O}$  is very efficient on cold silicate surface, Figure 8 shows the presence of only  $\sim 0.30 \text{ ML}$  of solid  $\text{D}_2\text{O}$  after the destruction of the oxygen ice by D atoms. This observation is well explained by the high desorption rate (88%) of water species. The most efficient pathway for water formation on the silicate substrate is through the formation of two OD radicals from Eq. (13) and the subsequent reaction (16) between OD and D atoms.

The reaction  $\text{D} + \text{OD} \rightarrow \text{D}_2\text{O}$  is highly exothermic and is then enable to release into the gas most ( $\sim 90\%$ ) of  $\text{D}_2\text{O}$  molecules upon its formation on the surface. The forma-

tion of  $\text{D}_2\text{O}$  water ice from  $\text{D} + \text{OD}$  reaction and through  $2\text{OD}$  as intermediates can explain our previous experimental results concerning the desorption of a large amount of  $\text{D}_2\text{O}$  molecules into the gas phase, mainly at the beginning ( $\sim 1000 \text{ s}$ ) of the D atoms-irradiation phase of  $\text{O}_2$  ice film (1 ML) (see Figure 6).

As D atoms react with oxygen molecules without an activation barrier and in the experiments on a silicate substrate, we have initially one monolayer of  $\text{O}_2$  on the sample, the formation of  $\text{D}_2$  is highly improbable. So, even if included in our reaction scheme (see Figure 7),  $\text{D} + \text{D}$  reaction does not affect the fits of  $\text{O}_2$ ,  $\text{D}_2\text{O}_2$ , and  $\text{D}_2\text{O}$  in Figure 8.

## V. ASTROPHYSICAL IMPLICATIONS

Our results have a great relevance in astrochemistry. We demonstrated for the first time from the experiments that water formation in the sub-monolayer and monolayer regimes can occur on bare amorphous silicates grains under interstellar conditions through surface reactions between atomic hydrogen and molecular oxygen. However, the desorption of species upon formation induced by the heat released from the exothermic reactions affects the efficiency of water formation on bare silicate grains at 10 K. Such a process revealed in this work is considered as the first evidence for the chemical desorption of water molecules in astrophysical environments where dust grains are composed of silicates. The result of our work shows that bare silicate surfaces have a large propensity to release species into the gas phase upon formation. In the ISM, the desorption of water molecules upon formation on dust grains is enhanced by the strong UV radiations field which dissociate the remaining molecules into radicals such as OH, O, and H. These newly intermediate species produce again water ices that desorb into the gas phase by the exothermic surface reactions. Following this desorption process of water, the formation of the first monolayer of water ice on bare grain surfaces such as amorphous silicates requires very long timescales to be grown. The time required to form the first layer of water ice in diffuse clouds is estimated to vary between ( $\sim 10^4$  years)<sup>26</sup> and ( $\sim 10^6$  years)<sup>7</sup> according to Monte Carlo simulations.

## VI. CONCLUSIONS

We have studied experimentally the initial stages of water formation through  $\text{O}_2 + \text{D}$  reaction on amorphous silicate surfaces analogous to bare dust grains in diffuse clouds. Co-deposition experiments of  $\text{O}_2$  and D atoms on the silicate substrate held at 10 K, in the condition  $[\text{D}]/[\text{O}_2] \sim 4$  have showed an effective formation of  $\text{D}_2\text{O}_2$  and  $\text{D}_2\text{O}$  molecules in the sub-monolayer and monolayer regimes. RAIRS results showed that the yield of  $\text{D}_2\text{O}$  water ice formed on the surface reaches  $\sim 20\%$ .

The kinetic studies performed by TPD experiments for the deuteration of one monolayer of  $\text{O}_2$  film, have revealed a H(D) isotopic effect for the formation of water through reaction (3) with activation barrier. The  $\text{D} + \text{D}_2\text{O}_2$  reaction occurring with tunneling process is much slower than  $\text{H} + \text{H}_2\text{O}_2$  reaction.

RAIRS and TPD experiments also showed that the efficiency of D<sub>2</sub>O water formation via Eley-Rideal mechanism between O<sub>2</sub> molecules adsorbed on the surface and D atom coming from the gas, depends strongly on the substrate whether it is composed of silicate or amorphous (H<sub>2</sub>O) water ice film analogous to dense molecular clouds. The fraction of D<sub>2</sub>O water ice formed on the surface of the silicate at 10 K was found to be ~20%, two and three times lower than the fractions of D<sub>2</sub>O water molecules formed on the porous amorphous water ice (~45%) and the nonporous amorphous water ice (~55%–60%) surfaces, respectively. It is possible that the catalytic effect of the water ice substrate enhances the reaction rates between D and O<sub>2</sub> by reducing the activation barriers and the exothermicity of the reactions.

A schematic reactions routes for water formation through O<sub>2</sub> + D pathway is proposed. It privileges the formation of two OD intermediates. Water ice is mainly formed on the silicate surface at 10 K through the barrierless and the exothermic D + OD → D<sub>2</sub>O reaction. Experiments have evidenced the desorption into the gas phase of more than (~85%) of D<sub>2</sub>O water molecules upon formation on the surface through the OD + D reaction. The high desorption rate of water molecules results from the excess energy released from the exothermic OD + D reaction. Such a chemical desorption of water ice from the silicates was supported by a kinetic model.

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- <sup>1</sup>E. L. Gibb, D. C. B. Whittet, A. C. A. Boogert, and A. G. G. M. Tielens, *Astrophys. J. Soc.* **151**, 35 (2004).
- <sup>2</sup>T. I. Hasegawa, E. Herbst, and C. M. Leung, *Astrophys. J. Soc.* **82**, 167 (1992).
- <sup>3</sup>A. G. G. M. Tielens and W. Hagen, *Astron. Astrophys.* **114**, 245 (1982).
- <sup>4</sup>D. Li, P. F. Goldsmith, and K. Menten, *Astrophys. J.* **587**, 262 (2003).
- <sup>5</sup>H. Mokrane, H. Chaabouni, M. Accolla, E. Congiu, F. Dulieu, M. Chehrouri, and J. L. Lemaire, *Astrophys. J. Lett.* **705**, L195 (2009).
- <sup>6</sup>D. Hollenbach, M. J. Kaufman, E. A. Bergin, and G. J. Melnick, *Astrophys. J.* **690**, 1497 (2009).
- <sup>7</sup>H. M. Cuppen and E. Herbst, *Astrophys. J. Lett.* **668**, 294 (2007).
- <sup>8</sup>C. Romanzin, S. Ioppolo, H. M. Cuppen, E. F. van Dishoeck, and H. Linnartz, *J. Chem. Phys.* **134**, 084504 (2011).
- <sup>9</sup>N. Miyauchi, H. Hidaka, T. Chigai, A. Nagaoka, N. Watanabe, and A. Kouchi, *Chem. Phys. Lett.* **456**, 27 (2008).
- <sup>10</sup>S. Ioppolo, H. M. Cuppen, C. Romanzin, E. F. van Dishoeck, and H. Linnartz, *Astrophys. J.* **686**, 1474 (2008).

- <sup>11</sup>E. Matar, E. Congiu, F. Dulieu, A. Momeni, and J. L. Lemaire, *Astron. Astrophys.* **492**, L17 (2008).
- <sup>12</sup>Y. Oba, N. Miyauchi, H. Hidaka, T. Chigai, N. Watanabe, and A. Kouchi, *Astrophys. J.* **701**, 464 (2009).
- <sup>13</sup>H. M. Cuppen, S. Ioppolo, C. Romanzin, and H. Linnartz, *Phys. Chem. Chem. Phys.* **12**, 12077 (2010).
- <sup>14</sup>S. Ioppolo, H. M. Cuppen, C. Romanzin, E. F. van Dishoeck, and H. Linnartz, *Phys. Chem. Chem. Phys.* **12**, 12065 (2010).
- <sup>15</sup>J. Troe and V. G. Ushakov, *J. Chem. Phys.* **128**, 204307 (2008).
- <sup>16</sup>R. Atkinson, D. L. Baulch, R. A. Cox, J. N. Crowley, R. F. Hampson, R. G. Hynes, M. E. Jenkin, M. J. Rossi, and J. Troe, *Atmos. Chem. Phys.* **4**, 1461 (2004).
- <sup>17</sup>Y. Oba, N. Watanabe, T. Hama, K. Kuwahata, H. Hidaka, and A. Kouchi, *Astrophys. J.* **749**, 67 (2012).
- <sup>18</sup>F. Dulieu, L. Amiaud, E. Congiu, J.-H. Fillion, E. Matar, A. Momeni, V. Pirronello, and J. L. Lemaire, *Astron. Astrophys.* **512**, A30 (2010).
- <sup>19</sup>D. Jing, J. He, J. Brucato, A. De Sio, L. Tozzetti, and G. Vidali, *Astrophys. J. Lett.* **741**, L9 (2011).
- <sup>20</sup>B. T. Draine, *Annu. Rev. Astron. Astrophys.* **41**, 241 (2003).
- <sup>21</sup>D. Williams and E. Herbst, *Surf. Sci.* **500**, 823 (2002).
- <sup>22</sup>F. Kemper, A. J. Markwick, and P. M. Woods, *Mon. Not. R. Astron. Soc.* **413**, 1192 (2011).
- <sup>23</sup>T. P. M. Goumans, C. R. A. Catlow, W. A. Brown, J. Kästner, and P. Sherwood, *Phys. Chem. Chem. Phys.* **11**, 5431 (2009).
- <sup>24</sup>F. Dulieu, E. Congiu, J. Noble, S. Baouche, H. Chaabouni, A. Moudens, M. Minissale, and S. Cazaux, "How micron-sized dust particles determine the chemistry of our Universe," *Sci. Rep.* (unpublished).
- <sup>25</sup>R. T. Garrod, V. Wakelam, and E. Herbst, *Astron. Astrophys.* **467**, 1103 (2007).
- <sup>26</sup>S. Cazaux, V. Cobut, M. Marseille, M. Spaans, and P. Caselli, *Astron. Astrophys.* **522**, A74 (2010).
- <sup>27</sup>E. A. Bergin, G. J. Melnick, J. R. Stauffer, M. L. N. Ashby, G. Chin, N. R. Erickson, P. F. Goldsmith, M. Harwit, J. E. Howe, S. C. Kleiner, D. G. Koch, D. A. Neufeld, B. M. Patten, R. Plume, R. Schieder, R. L. Snell, V. Tolls, Z. Wang, G. Winnewisser, and Y. F. Zhang, *Astrophys. J. Lett.* **539**, L129 (2000).
- <sup>28</sup>S. Baouche, Ph.D. dissertation, University of Cergy, 2004, see [http://www3.u-cergy.fr/LERMA-LAMAP/theses/Saoud\\_BAOUCHE\\_2004.pdf](http://www3.u-cergy.fr/LERMA-LAMAP/theses/Saoud_BAOUCHE_2004.pdf).
- <sup>29</sup>L. Amiaud, F. Dulieu, J.-H. Fillion, A. Momeni, and J. L. Lemaire, *J. Chem. Phys.* **127**, 144709 (2007).
- <sup>30</sup>J. L. Lemaire, G. Vidali, S. Baouche, M. Chehrouri, H. Chaabouni, and H. Mokrane, *Astrophys. J. Lett.* **725**, L156 (2010).
- <sup>31</sup>E. Congiu, H. Chaabouni, C. Laffon, P. Parent, S. Baouche, and F. Dulieu, *J. Chem. Phys.* **137**, 054713 (2012).
- <sup>32</sup>Z. Djouadi, L. D'Hendecourt, H. Leroux, A. P. Jones, J. Borg, D. Deboffle, and N. Chauvin, *Astron. Astrophys.* **440**, 179 (2005).
- <sup>33</sup>F. M. Hoffman, "Infrared reflection-absorption spectroscopy of adsorbed molecules," *Surf. Sci. Rep.* **3**, 107 (1983).
- <sup>34</sup>D. King and M. G. Wells, *Surf. Sci.* **29**, 454 (1972).
- <sup>35</sup>H. Chaabouni, H. Bergeron, S. Baouche, F. Dulieu, E. Matar, E. Congiu, L. Gavilan, and J. L. Lemaire, *Astron. Astrophys.* **538**, A128 (2012).
- <sup>36</sup>H. Uehara, K. Kawaguchi, and E. Hirota, *J. Chem. Phys.* **83**, 5479 (1985).
- <sup>37</sup>S. Hawkins, G. Kumi, S. Malyk, H. Reisler, and C. Wittig, *Chem. Phys. Lett.* **404**, 19 (2005).
- <sup>38</sup>M. Accolla, E. Congiu, F. Dulieu, G. Manicò, H. Chaabouni, E. Matar, H. Mokrane, J. L. Lemaire, and V. Pirronello, *Phys. Chem. Chem. Phys.* **13**, 8037 (2011).
- <sup>39</sup>M. Accolla, H. Chaabouni, E. Congiu, F. Dulieu, J. L. Lemaire, G. Manicò, and V. Pirronello, "Morphology of the solid water synthesized through the pathway D + O<sub>2</sub> studied by the sensitive TPD technique," *Mon. Not. R. Astron. Soc.* (submitted).
- <sup>40</sup>T. Zubkov, R. S. Smith, T. R. Engstrom, and B. D. Kay, *J. Chem. Phys.* **127**, 184707 (2007).
- <sup>41</sup>H. Hidaka, A. Kouchi, and N. Watanabe, *J. Chem. Phys.* **126**, 204707 (2007).
- <sup>42</sup>A. P. Jones and D. A. Williams, *Mon. Not. R. Astron. Soc.* **209**, 955 (1984).